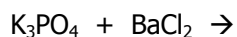


## TYPES OF REACTION REVIEW

1. Complete and balance the following equation:



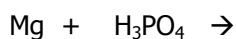
2. Complete the word equation, write the formula equation, and balance the following:



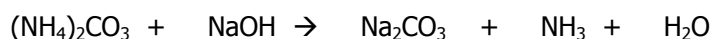
3. Balance the following equation:



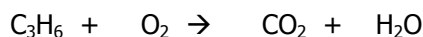
4. Complete and balance the following equation:



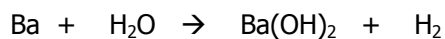
5. Balance the following equation:



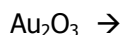
6. Balance the following equation, completing the equation first, if necessary.



7. Balance the following equation, completing the equation first, if necessary.



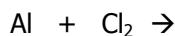
8. Balance the following equation, completing the equation first, if necessary.



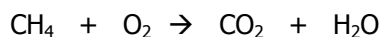
9. Balance the following equation, completing the equation first, if necessary.



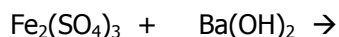
10. Balance the following equation, completing the equation first, if necessary.



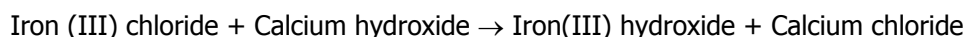
11. Balance the following equation, completing the equation first, if necessary.



12. Balance the following equation, completing the equation first, if necessary.



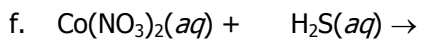
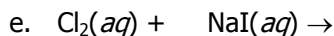
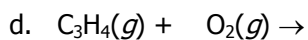
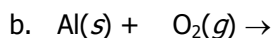
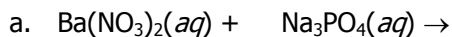
13. Rewrite the word equation as a balanced chemical equation.



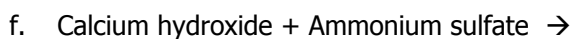
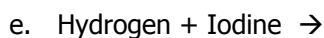
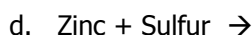
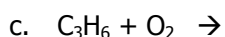
14. Write and balance an equation for the formation of the following compound from its elements.



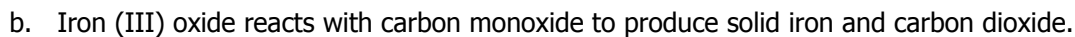
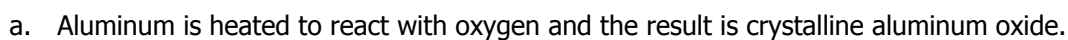
15. Complete and balance the following reactions, then identify the type of reaction it represents.



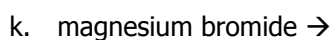
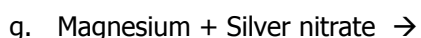
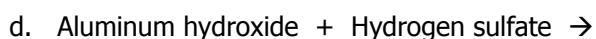
16. Complete, balance and identify the equations for each of the following reactions:



17. Write the complete chemical equation for each of the following reaction sentences (include reaction symbols where appropriate):



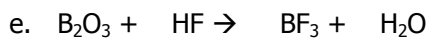
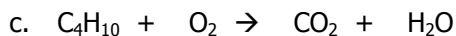
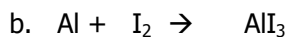
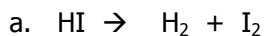
18. Complete and write the balanced formula equation for the following reactions, then identify the type of reaction each represents:



19. Write the balanced chemical equation for the following decomposition reaction.

Titanium (IV) hydroxide  $\rightarrow$  Titanium (IV) oxide + water

20. Balance the following chemical reactions:



21. Which of these compounds would be a precipitate as a product?

$\text{AgNO}_3$  or  $\text{Ag}_2\text{SO}_4$

22. Which of these compounds would be a precipitate as a product?

$\text{NH}_4\text{Cl}$  or  $(\text{NH}_4)_2\text{SO}_4$

23. Which of these compounds would be a precipitate as a product?

$\text{Ba}(\text{NO}_3)_2$  or  $\text{BaCO}_3$

24. Which of these compounds would be a precipitate as a product?

$\text{K}_2\text{SO}_4$  or  $\text{K}_2\text{CO}_3$

25. Chemicals to the left of the arrow in an equation are \_\_\_\_\_.

26. The arrow  $\uparrow$  means \_\_\_\_\_.

27. The catalyst is written where in an equation?

28. Combustion requires a hydrocarbon and \_\_\_\_\_ as the reactants.

29. The complete combustion of compounds containing carbon and hydrogen will produce \_\_\_\_\_ and \_\_\_\_\_ as the products.

30. The symbol that indicates that the compound is dissolved in water is \_\_\_\_\_.

31. A reaction in which simpler substances combine to form a more complex substance is called \_\_\_\_\_.

32. Would Copper replace Magnesium in a Magnesium sulfate solution?

33. Would Magnesium replace Copper in a Copper sulfate solution?