Periodic Table Chart Problem

I	II										III	IV	V	VI	VII	VIII
A																
B																
C	D				E			F			G		H	I		J
K	L			M					N		O	P	Q	R	S	
	T										U					V

1. Which of the elements has an electron configuration of s^2p^2 in the valence level? _____
2. Which of the elements in period 4 has the largest ionization energy? _____
3. Which of the elements in period 4 has the largest atomic radius? _____
4. Which of the elements in Group III has the highest density? _____
5. Which of the elements in Group II has the largest atomic radius? _____
6. Which of the elements in Group I has the highest ionization energy? _____
7. Which is the most metallic element of period 5? _____
8. Which of the elements in period 5 has the highest electronegativity? _____
9. Which of the elements are Noble Gases? _____
10. Which of the elements are Halogens? _____
11. Which of the elements are transition metals? _____
12. Which of the elements are Alkali Metals? _____
13. Which of the elements have two valence electrons? _____
14. Which of the elements are metalloids? _____
15. Which element has the lowest electronegativity? _____
16. Which of the elements tend to attract electrons to become negative ions? _____
17. Which of the elements tend to give away electrons to become positive ions? _____
18. Give the noble gas configuration for the element Q. _____
19. Give the noble gas configuration for the element C. _____
20. Give the noble gas configuration for the element S. _____
21. Which element in Group I will most likely have the highest melting point? _____
22. Which elements tend to give an ionic charge of +3? _____
23. Which elements tend to give an ionic charge of –3? _____
24. Which of the elements are nonmetals? _____
25. Which of the elements are metals? _____