

WORKSHEET ON SINGLE & DOUBLE REPLACEMENT REACTIONS

Predict the products. Write formulas & balance each reaction. If there is no reaction, then just put NO RXN.

Single Replacement: $A + BC \rightarrow B + AC$ or $A + BC \rightarrow C + BA$ (when A and C are negative ions)

1. Zinc + Hydrogen chloride \rightarrow
2. Magnesium + Hydrogen Sulfate \rightarrow
3. Copper (II) chloride + Fluorine \rightarrow
4. Silver + Sodium Hydroxide \rightarrow
5. Potassium iodide + Bromine \rightarrow
6. Calcium + Hydrogen hydroxide \rightarrow
7. Iron IV oxide + Hydrogen \rightarrow

Double Replacement: $AB + CD \rightarrow AD + CB$

1. Barium chloride + Aluminum sulfate \rightarrow
2. Calcium nitride + water \rightarrow
3. Calcium hydroxide + Hydrogen phosphate \rightarrow
4. Hydrogen sulfate + Sodium hydrogen carbonate \rightarrow
5. Calcium hydroxide + Ammonium chloride \rightarrow
6. Potassium iodide + Lead II Nitrate \rightarrow
7. Sodium acetate + Calcium sulfide \rightarrow

Complete each word equation, write formulas and balance the reaction equation. Then identify and place the type of reaction (single replacement or double replacement) in the blank provided.

1. Zinc + Silver nitrate \rightarrow
2. Aluminum + Hydrogen chloride \rightarrow
3. Magnesium oxalate + Ammonium carbonate \rightarrow
4. Calcium + Aluminum nitrate \rightarrow
5. Potassium fluoride + Lead (II) Nitrate \rightarrow
6. Calcium bromide + Silver nitrate \rightarrow
7. Ammonium phosphate + Barium acetate \rightarrow
8. Sodium chloride + Potassium \rightarrow
9. Magnesium nitrate + ammonium chloride \rightarrow
10. Iron (III) chlorate + calcium \rightarrow
11. Chlorine + Sodium bromide \rightarrow
12. Potassium chloride + Silver nitrate \rightarrow
13. Calcium hydroxide + Hydrogen nitrate \rightarrow
14. Lead II nitrate + Potassium chloride \rightarrow
15. Strontium carbonate + Hydrogen nitrate \rightarrow
16. Gold + Potassium nitrate \rightarrow
17. Zinc + Silver nitrate \rightarrow
18. Aluminum + Copper II sulfate \rightarrow